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a. $\text{Cu (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{CuO (s)}$
b. $\text{H}_2\text{O (l)} \rightarrow \text{H}_2 \text{ (g)} + \text{O}_2 \text{ (g)}$
c. $\text{Fe (s)} + \text{H}_2\text{O (g)} \rightarrow \text{H}_2 \text{ (g)} + \text{Fe}_3\text{O}_4 \text{ (s)}$
d. $\text{AsCl}_3 \text{ (aq)} + \text{H}_2\text{S (aq)} \rightarrow \text{As}_2\text{S}_3 \text{ (s)} + \text{HCl (aq)}$
e. $\text{CuSO}_4 \cdot 5 \text{H}_2\text{O (s)} \rightarrow \text{CuSO}_4 \text{ (s)} + \text{H}_2\text{O (l)}$

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Reaction Gezahegn Chaka, Ph. D.; Sudha Madhugiri, Ph. D. Collin College Chemistry Department
Objectives To calculate the thermodynamic quantities such as the changes in standard enthalpy, standard entropy, and standard free energy using equilibrium constant values calculated at various temperatures for a system of complexation reaction in aqueous solution.
Introduction Thermodynamic parameters are extr...

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